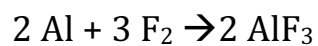
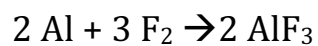


Station 2



- a) What is the molar mass of each of the compounds in the reaction above?
- b) Fluorine has a purity of 78%. How many grams of the product will be formed from 56.0 g of fluorine?

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Station 3

200.0 g of NaCl are dissolved in 100. mL of water. Calculate the molarity of the solution.

How many grams of AgCl are required to prepare 150.0 mL of 0.200 M solution?

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Station 4

Aluminum and hydrochloric acid react together to form hydrogen gas and aluminum chloride. What mass of AlCl_3 is produced when 24.5 g of Al reacts with 90.0 g of HCl?

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Station 5

In three trials of a titration, 36.9 mL, 34.4 mL, and 34.3 mL of 0.200 M NaOH were used to neutralize a 25.0 mL sample of H₂SO₄.

a) Write a balanced chemical reaction for this neutralization.

b) What was the average volume of NaOH used?

c) Calculate the molarity of the acid.

Station 5

In three trials of a titration, 36.9 mL, 34.4 mL, and 34.3 mL of 0.200 M NaOH were used to neutralize a 25.0 mL sample of H₂SO₄.

a) Write a balanced chemical reaction for this neutralization.

b) What was the average volume of NaOH used?

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Station 6

200.0 mL of 0.150 M AlCl_3 is added to 200.0 mL 0.250 M BaCl_2 . Calculate the $[\text{Ba}^{2+}]$, $[\text{Al}^{3+}]$ and the $[\text{Cl}^-]$ immediately after mixing the two solutions.

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