

# Activity: Building Covalent Compounds

Name:

Date:

Block:

## Part I: Review of Ionic Compounds

1. What type of elements (metals / non-metals) form an ionic compound?

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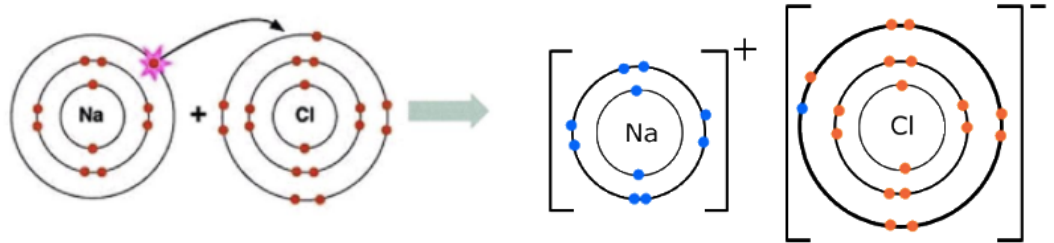
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2. Draw a Bohr model of the ionic compound lithium fluoride (remember, this is *after* electrons have been transferred and ions have formed)

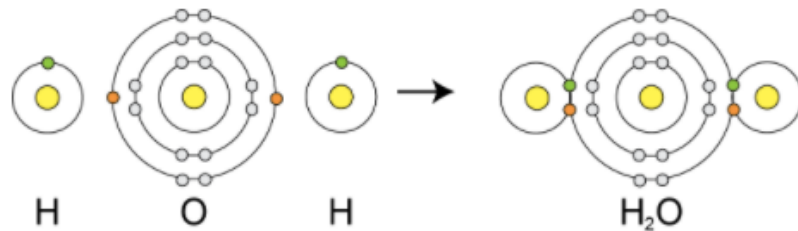
### LiF

3. What differences can you spot between the formation of an ionic compound versus a covalent compound?

Ionic Compound Formation:



Covalent Compound Formation:




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## Part II: Covalent Compounds

Using the provided materials, build covalent compounds and then draw their **Bohr models** on this sheet.

*Hint #1:* In covalent compounds, electrons are SHARED rather than transferred. No ions are made

*Hint #2:* Each spring represents two electrons that are shared between elements

*Hint #3:* Covalent compounds form individual molecules rather than a repeating lattice structure

Red: Oxygen

White: Hydrogen

Yellow: Sulfur

Green: Fluorine

Black: Carbon

Orange: Nitrogen

<b>H<sub>2</sub>O</b>	<b>HF</b>
<b>F<sub>2</sub></b>	<b>CH<sub>4</sub></b>
<b>H<sub>2</sub>S</b>	<b>NF<sub>3</sub></b>