

Chemistry V

Name:

Date:

Block:

1. Naming Ionic Compounds
2. Naming Covalent Compounds
3. Compounds with Multivalent Metals
4. Compounds with Polyatomic Ions

Naming Ionic Compounds

Review:

Ionic compounds consist of a metal and a non-metal ion.

- They occur when the metal transfers one or more electrons to the non-metal.
- They are bonded together by ionic bonds.

When elements form a compound, we are able to refer to these compounds using either its chemical name or its chemical formula.

Chemical Name:

The name of ionic compounds comes from the name of its elements. It is made up of two parts:

- 1.) The name of the METAL ion (positive ion) ALWAYS comes first.
 - a. We DO NOT change the name of the metal element.
- 2.) The name of the NON-METAL ion (negative ion) comes second.
 - a. We change the ending of the name of the non-metal ion to the suffix *-ide*.

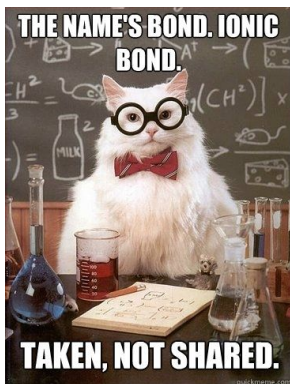


Table 2.5 Ions of Non-Metals

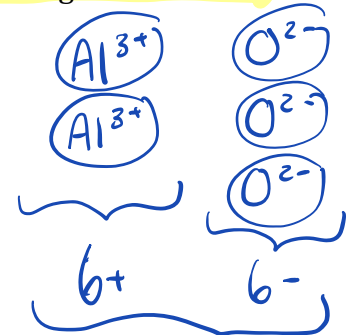
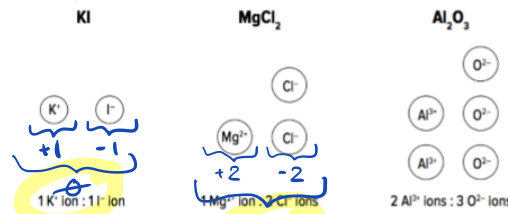
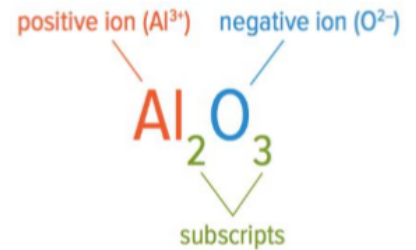
Element	Ion	Symbol	Group
fluorine	fluoride	F ⁻	17
chlorine	chloride	Cl ⁻	17
bromine	bromide	Br ⁻	17
iodine	iodide	I ⁻	17
oxygen	oxide	O ²⁻	16
sulfur	sulfide	S ²⁻	16
selenium	selenide	Se ²⁻	16
nitrogen	nitride	N ³⁻	15
phosphorus	phosphide	P ³⁻	15

Chemical Symbol	Chemical Name
metal → Li ₂ O ← nonmetal	Lithium oxide
CaF ₂	Calcium fluoride
MgS	Magnesium sulfide

Chemical Formula:

The chemical formula for a compound is composed of its chemical symbols.

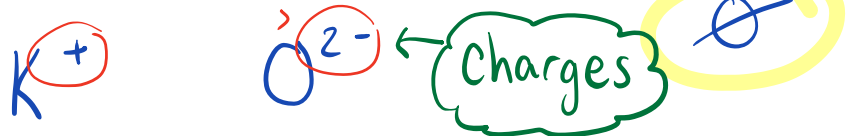
- The symbol for the **METAL ion** (positive ion) **ALWAYS** comes first.
- The symbol for the **NON-METAL ion** (negative ion) comes second.
- *Subscripts*** are used in order to indicate the ratio for each type of ion in the compound.
 - When there is **NO SUBSCRIPT**, we assume that the number is **1**.
- Although an ionic compound is made up of ions, the compound's **overall charge has to be zero** (positive charges must balance the negative charges).



How do we write the chemical formulas for ionic compounds?

- Identify and write the symbol and charge of the elements in the compound (Note: the **METAL** comes first!)

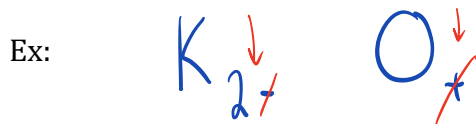
Ex: Potassium oxide



- SWAP:** Swap the charges of the ions

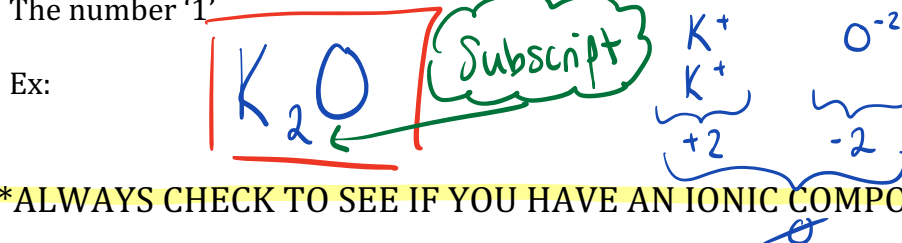


- DROP:** Drop the charges of the ions so they now become subscripts



- CHOP:** Take away the parts of the subscript that contain...

- The signs (+/-)
- If possible, **reduce the subscripts to the lowest terms** (i.e., Mn_2O_4 will become Mn_1O_2)
- The number '1'



*****ALWAYS CHECK TO SEE IF YOU HAVE AN IONIC COMPOUND!!!!*****

Chemical Name	Chemical Formula
Beryllium oxide	$\text{Be}^{+2} \text{O}^{-2} \rightarrow \text{Be}_1 \text{O}_1 \rightarrow \boxed{\text{BeO}}$
Scandium sulfide	$\text{Sc}^{+3} \text{S}^{-2} \rightarrow \text{Sc}_2 \text{S}_3 \rightarrow \boxed{\text{Sc}_2\text{S}_3}$
Sodium fluoride	$\text{Na}^{+} \text{F}^{-} \rightarrow \text{Na}_1 \text{F}_1 \rightarrow \boxed{\text{NaF}}$

Naming Binary Ionic Compounds

Name:

Date:

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1. Write the **chemical formula** for the following binary ionic compounds.

- | | |
|-----------------------|--|
| a. Sodium fluoride | $\text{Na}^+ \text{F}^- \rightarrow \boxed{\text{NaF}}$ |
| b. Potassium chloride | $\text{K}^+ \text{Cl}^- \rightarrow \boxed{\text{KCl}}$ |
| c. Calcium oxide | $\text{Ca}^{2+} \text{O}^{2-} \rightarrow \text{Ca}_1 \text{O}_1 \rightarrow \boxed{\text{CaO}}$ |
| d. Lithium bromide | $\text{Li}^+ \text{Br}^- \rightarrow \boxed{\text{LiBr}}$ |
| e. Beryllium chloride | $\text{Be}^{2+} \text{Cl}^- \rightarrow \boxed{\text{BeCl}_2}$ |
| f. Potassium sulfide | $\text{K}^+ \text{S}^{2-} \rightarrow \boxed{\text{K}_2\text{S}}$ |
| g. Magnesium nitride | $\text{Mg}^{2+} \text{N}^{3-} \rightarrow \boxed{\text{Mg}_3\text{N}_2}$ |
| h. Calcium phosphide | Ca_3P_2 |
| i. Magnesium oxide | MgO |
| j. Aluminum chloride | AlCl_3 |
| k. Sodium iodide | NaI |
| l. Potassium nitride | K_3N |
| m. Lithium sulfide | Li_2S |
| n. Barium phosphide | Ba_3P_2 |

2. Write the **name** for the following binary ionic compounds.

- | | |
|----------------------------|----------------------------|
| a. LiCl | <u>Lithium chloride</u> |
| b. MgS | <u>Magnesium sulfide</u> |
| c. NaCl | <u>Sodium chloride</u> |
| d. Al_2O_3 | <u>Aluminum oxide</u> |
| e. CaS | <u>Calcium sulfide</u> |
| f. Zn_3P_2 | <u>Zinc phosphide</u> |
| g. Be_3P_2 | <u>Beryllium phosphide</u> |
| h. Na_2O | <u>Sodium oxide</u> |
| i. KBr | <u>Potassium bromide</u> |
| j. BeO | <u>Beryllium oxide</u> |
| k. CaF_2 | <u>Calcium fluoride</u> |
| l. BaBr_2 | <u>Barium bromide</u> |
| m. Na_3N | <u>Sodium nitride</u> |
| n. Al_2S_3 | <u>Aluminum sulfide</u> |

Naming Covalent Compounds

Review:

Covalent compounds consist of only **non-metals**.

- They occur when the non-metals **share** their valence electrons with each other.
 - ***Note: these compounds DO NOT form ions.
- They are bonded together by covalent bonds.

Chemical Name:

The name of covalent compounds comes from the name of its elements. We use **prefixes** in order to indicate **how many atoms are present** in the molecule. *→ goes before the name*

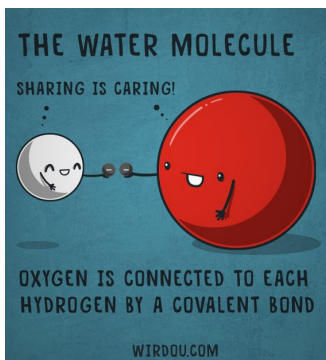


Table 2.8 Prefixes Used to Name Binary Covalent compounds

Prefix	Number	Prefix	Number
mono-	1	hexa-	6
di-	2	hepta-	7
tri-	3	octa-	8
tetra-	4	nona-	9
penta-	5	deca-	10

The name of a covalent compound is made up of two parts:

- 1.) Name the first non-metal element and add the appropriate **prefix IN FRONT** of the name.
 - a. We **DO NOT** use the prefix 'mono-' for the first element.
- 2.) Name the second non-metal element and change the ending to the suffix **-ide**. Add the appropriate prefix in front of the element's name.

***Note: When the end of the prefix and the beginning of the element's name contains the same letter (i.e., 'mono-' and 'oxide'), we can drop one of the repeating letters.

(no "mono" for first non-metal)

Chemical Symbol	Chemical Name
CO_2	Carbon dioxide
H_2O	dihydrogen monoxide
SF_6	sulfur hexafluoride

Chemical Formula:

The chemical formula for a compound is composed of its chemical symbols.

- The **prefixes** that are attached to the element's name will indicate the **subscript** for the element.
 - Keep in mind that these compounds do not form ions. You do not need to look at the ion charges!
- For the first element, if it does not have a prefix in its name, we assume that the subscript is 1.

***Note: We **DO NOT reduce** the subscripts for covalent compounds. When there is no subscript, we assume that the number is 1.

***Note: **DO NOT do the drop, swap, and chop** method for covalent compounds. We only have to look at the prefixes of the compound's name to find the chemical formula.

Chemical Name	Chemical Formula
<u>D</u> initrogen <u>t</u> etroxide	N_2O_4 * Don't reduce!
Nitrogen <u>m</u> onoxide	NO
Phosphorus <u>t</u> etrahydride	PH_4

ALWAYS CHECK TO SEE IF YOU HAVE A COVALENT COMPOUND!!!!

Practice:

1. Write the **chemical formula** for the following binary covalent compounds.

- Sulfur dioxide SO_2
- Carbon tetrafluoride CF_4
- Selenium trioxide SeO_3
- Nitrogen trichloride NCl_3
- Carbon dioxide CO_2
- Boron trifluoride BF_3
- Tetrasulfur tetranitride S_4N_4
- Diphosphorus pentoxide P_2O_5
- Carbon disulfide CS_2
- Nitrogen monoxide NO
- Diasrsenic trioxide As_2O_3
- Sulfur hexafluoride SF_6

2. Write the **name** for the following binary covalent compounds.

- Cl_2O_7 dichlorine heptoxide
- N_2O_4 dinitrogen tetroxide
- SO_3 sulfur trioxide
- PCl_5 phosphorus pentachloride
- NF_3 nitrogen trifluoride
- CS_2 Carbon disulfide
- SiF_4 Silicon tetrafluoride
- N_2O_3 dinitrogen trioxide
- BI_3 boron triiodide
- ClF_3 Chlorine trichloride
- SCl_2 Sulfur dichloride
- CO Carbon monoxide

Compounds with Multivalent Metals

A **multivalent metal** is a metal element that is able to form **two or more types of ions** with different charges. We tend to find the multivalent metals in families **3 - 12**.

In order to indicate which charge the metal ion forms, we use **roman numerals** right **after the name**.

- Example: Copper can form ions with a +1 or a +2 charge.
 - Cu^+ : Copper (I)
 - Cu^{+2} : Copper (II)

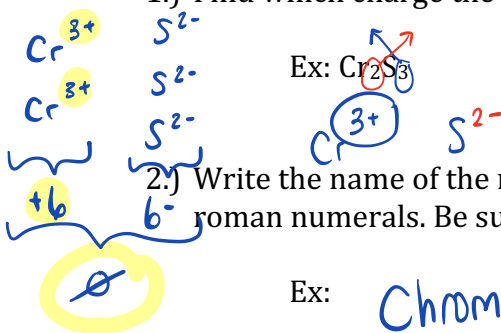
Table 2.6 Roman Numerals

Metal Ion Charge	Roman Numeral
1+	I
2+	II
3+	III
4+	IV
5+	V
6+	VI
7+	VII

Chemical Name:

Multivalent metals will always form an ionic compound. In order to find the chemical name of the compound, we have to first find out which charge the multivalent metal ion will form.

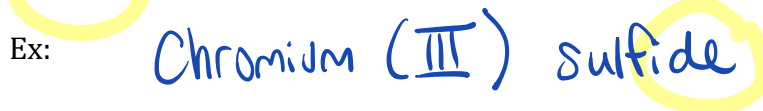
- 1.) Find which charge the multivalent metal ion will form. "reverse swap & drop"



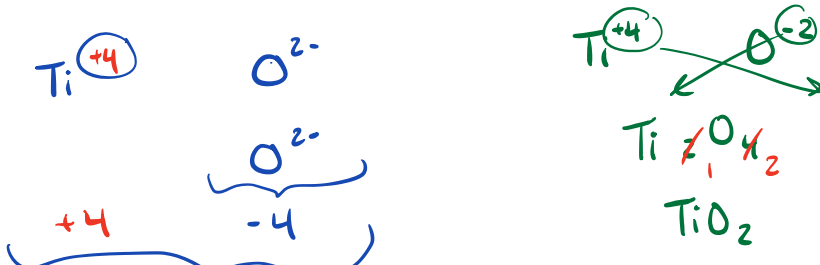
- 2.) Write the name of the metal ion (positive ion) first and indicate the charge of the ion with roman numerals. Be sure to place the roman numerals in brackets after the element's name.



- 3.) Write the **name** of the non-metal ion (negative ion) after the roman numerals and change the ending to '-ide'.



Chemical Symbol	Chemical Name
Fe_2O_3	Iron (III) oxide
CoF_2	Cobalt (II) fluoride
TiO_2	Titanium (IV) oxide

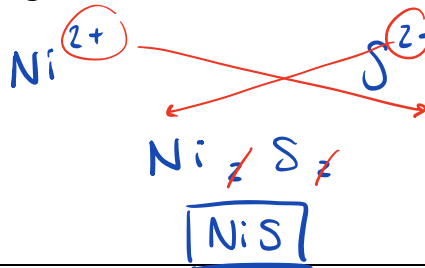


Chemical Formula:

The chemical formula for compounds that contain multivalent metals can be found by using the **SWAP, DROP, and CHOP** method as ALL multivalent metals will form an **ionic compound**.

- The roman numeral in the chemical name will indicate the ion charge of the multivalent metal. This will always be a positive charge.

Ex: Nickel (II) Sulfide



Chemical Name	Chemical Formula
Chromium (III) chloride	$\text{Cr}^{3+} \text{Cl}^{-} \rightarrow \boxed{\text{CrCl}_3}$
Cobalt (II) sulfide	$\text{Co}^{2+} \text{S}^{2-} \rightarrow \text{Co}_1 \text{S}_1 \rightarrow \boxed{\text{CoS}}$
Manganese (IV) bromide	$\text{Mn}^{4+} \text{Br}^{-} \rightarrow \boxed{\text{MnBr}_4}$

Practice:

- Write the formulas and names of the ionic compounds with the following combination of ions. The table has been partially completed to help guide you

Positive Ion	Negative Ion	Chemical Formula	Compound Name
Ti^{3+}	Cl^{-}	TiCl_3	Titanium (III) chloride
Fe^{2+}	O^{2-}	FeO	Iron (II) oxide
Pd^{2+}	Br^{-}	PdBr_2	Palladium (II) bromide
Sn^{4+}	F^{-}	SnF_4	Tin (IV) fluoride
Au^{+}	Cl^{-}	AuCl	Gold (I) chloride
Pt^{4+}	O^{2-}	PtO_2	Platinum (IV) oxide
Co^{2+}	F^{-}	CoF_2	Cobalt (II) fluoride
Ni^{2+}	I^{-}	NiI_2	Nickel (II) iodide
Nb^{3+}	N^{3-}	NbN	Niobium (III) nitride
Mn^{4+}	O^{2-}	MnO_2	Manganese (IV) oxide

- Write the chemical formulas of the following ionic compounds

Cobalt (III) fluoride CoF_3
 Osmium (IV) chloride OsCl_4
 Chromium (III) oxide Cr_2O_3
 Mercury (II) selenide $\text{Hg}_2\text{Se}_2 \rightarrow \text{HgSe}$
 Copper (II) chloride CuCl_2

Lead (II) sulfide PbS
 Titanium (III) nitride TiN
 Bismuth (III) sulfide Bi_2S_3
 Ruthenium (IV) oxide $\text{Ru}_2\text{O}_4 \rightarrow \text{RuO}_2$
 Nickel (II) fluoride NiF_2

- Write the names of the following ionic compounds

NbCl_5 Niobium (V) chloride
 SnF_4 Tin (IV) fluoride
 Mn_2O_3 Manganese (III) oxide
 RhCl_3 Rhodium (III) chloride
 NiF_3 Nickel (III) fluoride

* HgS Mercury (II) sulfide
 TlI Thallium (I) iodide
 * IrO_2 Iridium (IV) oxide
 FeCl_2 Iron (II) chloride
 V_2O_5 Vanadium (V) oxide

Compounds with Polyatomic Ions

many atoms

A polyatomic ion is an ion made up of **two or more covalently** bonded atoms.

- Example: carbonate ion (CO_3^{2-})
 - 1 carbon atom
 - 3 oxygen atoms

↳ all non-metals
(some exceptions - Cr)
Mn

Polyatomic ions will form an ion charge. If the ion charge is positive, we will treat the polyatomic ion as a metal. If the ion charge is negative, we will treat the polyatomic ion as a non-metal.

Table 2.7 Names, formulas, and charges of some common polyatomic ions

1+ Charge	1- Charge	2- Charge	3- Charge
ammonium, NH_4^+	acetate, CH_3COO^-	carbonate, CO_3^{2-}	phosphate, PO_4^{3-}
	chlorate, ClO_3^-	chromate, CrO_4^{2-}	phosphite, PO_3^{3-}
	chlorite, ClO_2^-	dichromate, $\text{Cr}_2\text{O}_7^{2-}$	
	hydrogen carbonate, HCO_3^-	peroxide, O_2^{2-}	
	hydroxide, OH^-	sulfate, SO_4^{2-}	
	nitrate, NO_3^-	sulfite, SO_3^{2-}	
	nitrite, NO_2^-		
	permanganate, MnO_4^-		

We can treat polyatomic ions as **one large element**. We CANNOT separate polyatomic ions into its parts.

Chemical Name:

Polyatomic ions will always form an **ionic compound**. We **DO NOT** change the name of the polyatomic ion (i.e., we do not change the suffix of the polyatomic ion to '-ide').

Ex: $\text{Ca}_3(\text{PO}_4)_2$

Calcium phosphate

↑ double check for multivalent

Chemical Symbol	Chemical Name
$\text{Ca}(\text{NO}_3)_2$	Calcium nitrate
$\text{Be}(\text{ClO})_2$	Beryllium hypochlorite
$(\text{NH}_4)_2\text{S}$	Ammonium sulfide

↑ change to -ide

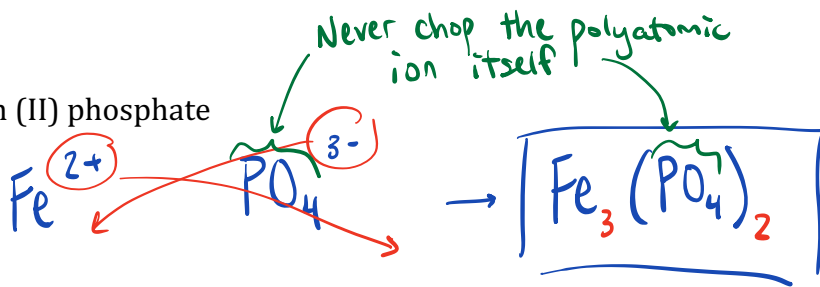
Chemical Formula:

The chemical formula for compounds that contain polyatomic ions can be found by using the **SWAP**, **DROP**, and **CHOP** method as ALL compounds that have a polyatomic ion will form an ionic compound.

- When we have more than one group of polyatomic ions in a compound, we will have to place **brackets** around the polyatomic ion symbols before writing the subscript.
 - Be sure to place the subscript **outside of the brackets**.

***Note: The subscripts that are contained within the polyatomic ion are NOT part of the 'chop' step. We DO NOT change the symbols or the subscripts of the polyatomic ion.

Ex: Iron (II) phosphate



Chemical Name	Chemical Formula
Calcium nitrate	$\text{Ca}^{2+} \text{NO}_3^- \rightarrow \text{Ca}(\text{NO}_3)_2$
Sodium hydroxide	$\text{Na}^+ \text{OH}^- \rightarrow \text{NaOH}$ * no bracket needed b/c no subscript with OH
Cobalt (III) sulfite	$\text{Co}^{3+} \text{SO}_3^{2-} \rightarrow \text{Co}_2(\text{SO}_3)_3$

Practice:

1. Complete the table by providing the correct chemical formula for each compound formed from the ions indicated

	F^- Fluoride	OH^- Hydroxide	CO_3^{2-} Carbonate	PO_4^{3-} Phosphate
Na^+ Sodium	NaF	NaOH	Na_2CO_3	Na_3PO_4
Al^{3+} Aluminum	AlF_3	$\text{Al}(\text{OH})_3$	$\text{Al}_2(\text{CO}_3)_3$	AlCO_3
Cu^{2+} Copper (II)	CuF_2	$\text{Cu}(\text{OH})_2$	CuCO_3	$\text{Cu}_3(\text{PO}_4)_2$
Mn^{4+} Manganese (IV)	MnF_4	$\text{Mn}(\text{OH})_4$	$\text{Mn}(\text{CO}_3)_2$	$\text{Mn}_3(\text{PO}_4)_4$
NH_4^+ Ammonium	NH_4F	NH_4OH	$(\text{NH}_4)_2\text{CO}_3$	$(\text{NH}_4)_3\text{PO}_4$

2. Write the chemical formulas of the following ionic compounds

Barium chlorite $\text{Ba}(\text{ClO}_2)_2$
 Nickel (II) nitrate $\text{Ni}(\text{NO}_3)_2$
 Potassium chromate K_2CrO_4
 Lead (IV) phosphate $\text{Pb}_3(\text{PO}_4)_4$

Copper (I) carbonate Cu_2CO_3
 Chromium (III) sulfite $\text{Cr}_2(\text{SO}_3)_3$
 Calcium phosphite $\text{Ca}_3(\text{PO}_3)_2$
 Iron (III) acetate $\text{Fe}(\text{CH}_3\text{COO})_3$
 Strontium permanganate $\text{Sr}(\text{MnO}_4)_2$

3. Write the names of the following ionic compounds

$\text{V}(\text{ClO}_3)_4$ Vanadium (IV) chlorate
 $\text{Al}_2(\text{CO}_3)_3$ Aluminum carbonate
 $\text{Co}(\text{NO}_2)_2$ Cobalt (II) nitrite
 $(\text{NH}_4)_2\text{SO}_4$ Ammonium sulfate
 * $\text{Ti}(\text{CrO}_4)_2$ Titanium (IV) chromate

Li_3PO_4 Lithium phosphate
 $\text{Cr}(\text{MnO}_4)_3$ Chromium (III) permanganate
 $\text{Ag}_2\text{Cr}_2\text{O}_7$ Silver dichromate
 $\text{Mg}(\text{HCO}_3)_2$ Magnesium bicarbonate
 $\text{Sn}(\text{OH})_4$ Tin (IV) hydroxide