

Chemistry 12

Acid-Base Equilibrium V Check Point

✓ Hydrolysis

Name:

Date:

Block:

Calculate the pH of the solution produced when 16.0 g of ammonium nitrate is dissolved in enough water to produce 500.0 mL of solution.

The pH of a 0.50 M solution of $\text{N}_2\text{H}_5\text{Cl}$ is found to be 4.266. Calculate the K_a for N_2H_5^+ .

Calculate the pH of a 0.10M $(\text{NH}_4)_2\text{HPO}_4$ solution.