

Periodic Trends Worksheet 2

Name:

Date:

Block:

1. Arrange the following from largest atomic radius to smallest atomic radius:

a) Po, O, Se, S



b) W, Hg, Cs, La



2. Arrange the following from highest ionization energy to lowest ionization energy:

a) At, F, Br, Cl



b) Sb, Rb, Ag, I



3. Arrange the following from most electronegative to least electronegative:

a) Bi, N, As, P, Sb



b) Sc, Ge, Cu, Ti, Kr

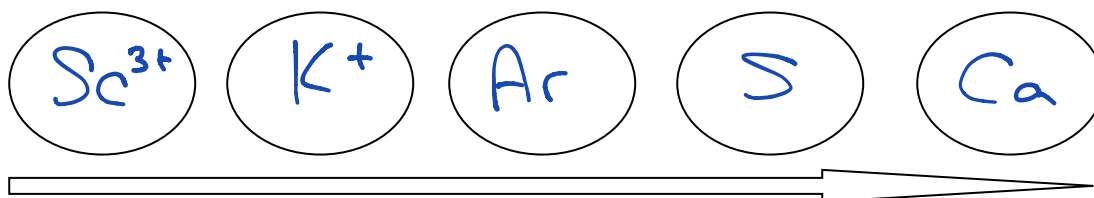


4. Use the following particles to answer the questions below:

Chemical Name:	18p 18e ⁻ Argon atom	21p 18e ⁻ Scandium ion	16p 16e ⁻ Sulfur atom	20p 20e ⁻ Calcium atom	19p 18e ⁻ Potassium ion
Chemical Formula:	Ar	Sc ³⁺	S	Ca	K ⁺

has more shells = biggest

a. Rank the particles in atomic size from smallest to biggest (you must use **chemical formulas** and *not* the chemical name):



b. Which of the particles would have the greatest electronegativity?



c. Which of the particles would have the greatest ionization energy?

