

Periodic Trends Worksheet 2

Name:

Date:

Block:

1. Arrange the following from largest atomic radius to smallest atomic radius:

a) Po, O, Se, S

b) W, Hg, Cs, La

2. Arrange the following from highest ionization energy to lowest ionization energy:

a) At, F, Br, Cl

b) Sb, Rb, Ag, I

3. Arrange the following from most electronegative to least electronegative:

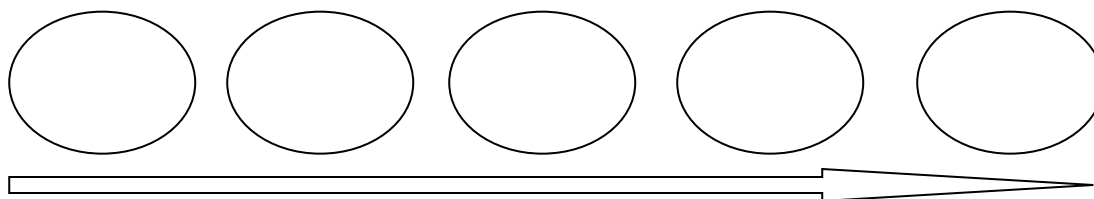
a) Bi, N, As, P, Sb

b) Sc, Ge, Cu, Ti, Kr

4. Use the following particles to answer the questions below:

Chemical Name:	Argon atom	Scandium ion	Sulfur atom	Calcium atom	Potassium ion
Chemical Formula:					

a. Rank the particles in atomic size from smallest to biggest (you must use **chemical formulas** and *not* the chemical name):



b. Which of the particles would have the greatest electronegativity? _____

c. Which of the particles would have the greatest ionization energy? _____