

**Lab: Electrolytic Cells**

Name: \_\_\_\_\_

Block: \_\_\_\_\_

**For Students:**

Lab performed: \_\_\_\_\_

Lab due: \_\_\_\_\_

**For Teacher:**Pre-lab completion:  Yes  NoLab Submitted:  On Time  Late**Introduction:**

In electrolytic cells, \_\_\_\_\_ produces a \_\_\_\_\_.

**Fill in the blanks:**

Half Reaction:

Half Reaction:

Observations:

Observations:

Total voltage required: \_\_\_\_\_

**Objectives:**

- 1.
- 2.
- 3.

**Procedure:**

Part I: Electrolysis of 1.0M ZnSO<sub>4</sub> (aq)

**Data & Observations:**

<u>Electrode</u>	<u>Observations</u>	<u>Half-Reactions</u>
Anode		
Cathode		

*Overall Reaction:*

*Overall Voltage:*

**Procedure:**

Part II: Copper Plating

**Data & Observations:**

Mass (g) of object to be plated:

Before: \_\_\_\_\_

After: \_\_\_\_\_

Difference: \_\_\_\_\_

<u>Electrode</u>	<u>Observations</u>	<u>Half-Reactions</u>
Anode		
Cathode		

*Overall Reaction:*

*Overall Voltage:*

**Follow Up Questions:**

Draw and label the electrolytic cells for Part I and Part II.

Part I:

Part II: