Chemistry 11

Lab: Precipitate Formation

Name:
Date:
Block:

Objectives:

1. To mix several pairs of solutions together and then note whether any precipitates form.
2. To deduce, from experimental results, which combinations of ions form precipitates.
3. To write a balanced formula equation, complete ionic equation and net ionic equation for each precipitation reaction.

Safety:

- Wash any spills and splashes immediately with plenty of water.
- Notify your instructor.

Procedure:

1. Place a drop of one solution on the spot plate and add a drop of a second solution to it. Your observation will be enhanced if you place the glass square over a dark surface.
2. If a precipitate forms (a solid), record this result in the table below by writing "**ppt**" in the appropriate square, along with the **colour** of the precipitate. If no precipitate forms, simply mark a dash (–) in the square.
3. Repeat the above steps until all possible combinations of solutions have been tested.
4. Rinse out all equipment and dry the glass square with a paper towel.
5. Wash your hands thoroughly with soap and water.

Analysis:

For each different precipitation, write the following:

a) Balanced formula equation
b) Complete net ionic equation
c) Net ionic equation
d) Identify the solid that was formed
e) Identify the spectators ions

Solutions I:

Fill in the grey boxes of the table with the IONS from each solution that are available to you.

Solutions Used						

How many precipitates did you observe? _____

Analysis:

Solutions II:

Fill in the grey boxes of the table with the IONS from each solution that are available to you.

Solutions Used						

How many precipitates did you observe? _____

Analysis:

Solutions III:

Fill in the grey boxes of the table with the IONS from each solution that are available to you.

Solutions Used						

How many precipitates did you observe? _____

Analysis: