

Station 1:

It is found that 1.892×10^{-13} grams of the compound cadmium (II) sulphide will dissolve in 350.0 mL of water to form a saturated solution. Using this data, calculate the value for the K_{sp} of CdS

Station 2:

Up to 15.0g of barium chloride can be dissolved in 2.5L of $Al_2(SO_4)_3$ solution without forming a precipitate. Find the mass of aluminum in the solution.

How did you do?

<i>"I don't get it yet, but I'm trying."</i>	<i>"I'm starting to get it."</i>	<i>"I get it."</i>	<i>"I really get it and can teach others how to do it."</i>
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Station 3:

Calculate the mass of Na_2CO_3 that must be added to 2.50 L of 0.00085 M MgCl_2 in order to just start precipitation

Station 4:

A sample of a saturated solution of MgF_2 was evaporated and the following data table was constructed:

Mass of empty evaporating dish: 78.5418 g
Mass of evaporating dish and MgF_2 residue after evaporation: 78.5434 g
Volume of saturated MgF_2 : 100.00 mL
Temperature: 25.0 °C

Use this data to calculate the value of K_{sp} for MgF_2 at 25°C

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Station 5:

What is the maximum mass of copper (II) chloride you can add 100.0L of a 0.025M solution of sodium iodate without causing precipitation?

Station 6:

What is the maximum volume of 0.0350M sodium sulphate solution required to obtain a saturated solution of strontium sulphate with 1.25g of strontium nitrate?

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Station 7:

Calculate the $[Ag^+]$ required to just start precipitation of Ag_2CO_3 in a 0.0030 M solution of $(NH_4)_2CO_3$

Station 8:

If 250.0 mL of 3.40×10^{-4} M $Cu(NO_3)_2$ is mixed with 350.0 mL of 3.12×10^{-4} M KIO_3 , will a precipitate form?

How did you do?

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