Chemistry 11

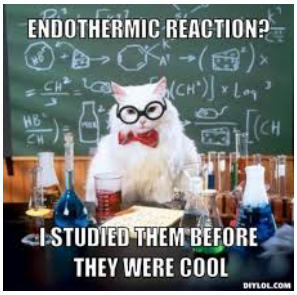
Stoichiometry VI

Name:
Date:
Block:

1. Enthalpy (Heat Energy)

Enthalpy (Heat Energy)

- Systems can either gain energy or release energy, often felt as heat
- Energy is measured as _____
- Symbol = _____
- Unit = _____ (_____)



Gain Energy

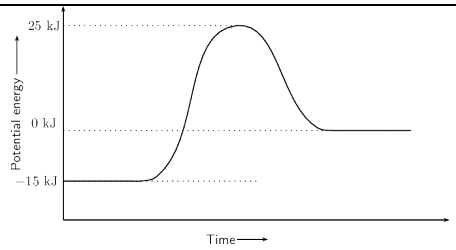
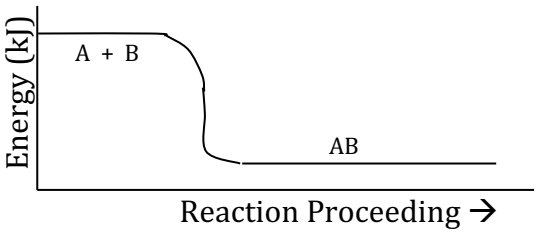
- _____ thermic
- ΔH =
-
- System feels _____
- _____ favoured.

Release Energy

- _____ thermic
- ΔH =
-
- System feels _____
- _____ favoured.

Note: Systems favour _____ energy states (i.e. _____ enthalpy)

Practice:

Reaction:	Is the reaction endothermic or exothermic?	Which side is favoured?
1. K + D + heat → G		
2. U → C ΔH = - 60 kJ		
3. F + A → T + I ΔH = 60 kJ		
4. A + B → C + heat		
5. 		
6. H_2 + Cl_2 → 2 HCl + 432 kJ		
7. $12CO_2$ + $11H_2O$ → $C_{12}H_{22}O_{11}$ + 12 O_2 ΔH = 5638 kJ		
8. 		
9. C + D → CD ΔH= -65.7 kJ		
10. E + F + 437 kJ → G + H		